Department: Science Chemistry

Year 9

Plan Of Learning For The Year (Unit/Topic/Project Context)			
Half Term 1	Half Term 4		
Chemistry - Atomic Structure	Chemistry – Analysis		
Half Term 2	Half Term 5		
Chemistry – Atomic Structure and Bonding	Chemistry - Analysis		
Half Term 3	Half Term 6		
Chemistry – Bonding	 Review of the Year to include end of year assessment and feedback 		

Feedback, Retrieval & Assessment	Super curriculum opportunities / extra-curricular activities	Cultural Capital, SMSC, Careers and Futures
 Regular self and peer assessment Regular Formative Assessment Termly Teacher Assessment (FFA) Termly Formal Assessment (FA) Whole Class Feedback 	 Use of Seneca to Support Learning both as homework and independent study 	 Career Link in Each Unit, linking to the Gatsby Benchmark Opportunities for Practical work that both links and applies to industry Transferable skills via practicals such as problem solving, group work and working to a deadline.

Common misconceptions	Connecting New Knowledge	Challenge for all
 Atoms do not weigh anything Relative particle spacing among solids, liquids, and gases is incorrectly perceived and not generally related to the densities of the states. (Microscopic model does not represent macroscopic properties.) Materials can only exhibit properties of one state of matter. Particles possess the same properties as the materials they compose. For example, atoms of copper are "orange and shiny," gas molecules are "transparent," and solid molecules are "Hard." 	 Regular revisiting of core vocabulary and key concepts, building key knowledge for GCSE Provision of Knowledge Organisers for each topic given at the start Provision of Curriculum map 	 Regular use of scaffolds and structured practice Clearly defined success criteria and use of clear feedback model to show next steps to improve Stretch activates built in to each lesson